

REGULAR Common Ions

CATIONS (+)			ANIONS (-)		
Name	Symbol	Alternative*	Name	Symbol	Alternative*
Aluminum	Al ³⁺		Bromide	Br ⁻	
Ammonium	NH ₄ ⁺		Bromate (I)	BrO ⁻	(Hypobromite)
Arsenic (III)	As ³⁺		Bromate (III)	BrO ₂ ⁻	(Bromite)
Arsenic (V)	As ⁵⁺		Bromate (V)	BrO ₃ ⁻	(Bromate)
Barium	Ba ²⁺		Bromate (VII)	BrO ₄ ⁻	(Perbromate)
Bismuth (III)	Bi ³⁺		Carbonate	CO ₃ ²⁻	
Bismuth (V)	Bi ⁵⁺		Chlorate (I)	ClO ⁻	(Hypochlorite)
Cadmium	Cd ²⁺		Chlorate (III)	ClO ₂ ⁻	(Chlorite)
Calcium	Ca ²⁺		Chlorate (V)	ClO ₃ ⁻	(Chlorate)
Chromium (II)	Cr ²⁺		Chlorate (VII)	ClO ₄ ⁻	(Perchlorate)
Chromium (III)	Cr ³⁺		Chloride	Cl ⁻	
Cobalt (II)	Co ²⁺		Chromate	CrO ₄ ²⁻	
Cobalt (III)	Co ³⁺		Cyanide	CN ⁻	
Copper (I)	Cu ⁺	(Cuprous)	Dichromate	Cr ₂ O ₇ ²⁻	
Copper (II)	Cu ²⁺	(Cupric)	Dihydrogen Phosphate	H ₂ PO ₄ ⁻	
Hydrogen	H ⁺		Ethanoate	C ₂ H ₃ O ₂ ⁻	(Acetate)
Hydronium	H ₃ O ⁺		Fluoride	F ⁻	
Iron (II)	Fe ²⁺	(Ferrous)	Hydride	H ⁻	
Iron (III)	Fe ³⁺	(Ferric)	Hydrogen Carbonate	HCO ₃ ⁻	(Bicarbonate)
Lead (II)	Pb ²⁺	(Plumbous)	Hydrogen Oxalate	HC ₂ O ₄ ⁻	(Binoxalate)
Lead (IV)	Pb ⁴⁺	(Plumbic)	Hydrogen Phosphate	HPO ₄ ²⁻	
Lithium	Li ⁺		Hydrogen Sulfate	HSO ₄ ⁻	(Bisulfate)
Magnesium	Mg ²⁺		Hydrogen Sulfide	HS ⁻	(Bisulfide)
Manganese (II)	Mn ²⁺		Hydrogen Sulfite	HSO ₃ ⁻	(Bisulfite)
Manganese (IV)	Mn ⁴⁺		Hydroxide	OH ⁻	
Mercury (I)	Hg ₂ ⁺	(Mercurous)	Iodate (I)	IO ⁻	(Hypoiodite)
Mercury (II)	Hg ²⁺	(Mercuric)	Iodate (III)	IO ₂ ⁻	(Iodite)
Nickel (II)	Ni ²⁺		Iodate (V)	IO ₃ ⁻	(Iodate)
Potassium	K ⁺		Iodate (VII)	IO ₄ ⁻	(Periodate)
Silver	Ag ⁺		Iodide	I ⁻	
Sodium	Na ⁺		Manganate (VII)	MnO ₄ ⁻	(Permanganate)
Strontium	Sr ²⁺		Nitrate	NO ₃ ⁻	
Tin (II)	Sn ²⁺	(Stannous)	Nitride	N ³⁻	
Tin (IV)	Sn ⁴⁺	(Stannic)	Nitrite	NO ₂ ⁻	
Zinc	Zn ²⁺		Oxalate	C ₂ O ₄ ²⁻	(Ethandioate)
			Oxide	O ²⁻	
			Peroxide	O ₂ ²⁻	
			Phosphate	PO ₄ ³⁻	
			Phosphide	P ³⁻	
			Phosphite	PO ₃ ³⁻	
			Sulfate	SO ₄ ²⁻	
			Sulfide	S ²⁻	
			Sulfite	SO ₃ ²⁻	
			Thiosulfate	S ₂ O ₃ ²⁻	
			Thiocyanate	SCN ⁻	

* For cations, the alternative names are seldom used in modern chemistry. For the anions, however, the alternatives are sometimes used. For example, the oxyhalogen ions like bromate, chlorate and iodate are usually referred to by the alternate names.