

Rules for assigning oxidation numbers

Elemental form	zero (0) . Only one kind of atom present, no charge
Atomic ions	= the charge on the atom (monatomic ion)
Group 1A Li,Na,K,Rb,Cs	+1 unless in elemental form
Group 2A Be,Mg,Ca,Sr,Ba	+2 unless in elemental form
Hydrogen (H)	+1 when bonded to a nonmetal, -1 when bonded to a metal
Oxygen (O)	-1 in peroxides O_2^- , -2 in all other compounds (most common)
Fluorine (F)	-1 , always
Neutral compounds	The sum of all oxidation numbers of atoms or ions in a neutral compound is zero .
Ionic compounds	The sum of all oxidation numbers of atoms in an ionic compound is the charge on the polyatomic ion.

Common Oxidation Numbers of the First 39 Elements

